

Walter A. Hynes, Leo K. Yanowski and Morris Schiller. A Modified Method for the Preparation of Monochloro-pentamminocobaltic Chloride (Purpureocobaltic Chloride).

Page 3053. "The method described here duplicates almost exactly one proposed by Willard and Hall in their paper, 'The Separation of Cobalt by means of Phenylthiohydantoic Acid,' THIS JOURNAL, 44, 2219 (1922). Since the method of Dr. Willard was mentioned as a mere incidental to the balance of the work reported and was not indicated in the published title of his article, we had not included his article in our list of literature references."—W. A. HYNES.

George Scatchard, W. J. Hamer and S. E. Wood. Isotonic Solutions. I. The Chemical Potential of Water in Aqueous Solutions of Sodium Chloride, Potassium Chloride, Sulfuric Acid, Sucrose, Urea and Glycerol at 25°.

Page 3066. The points at concentrations above 1 *M* representing the measurements of Frazer and Norris in Fig. 1 and those of Lovelace, Frazer and Sease in Fig. 2, should be lowered about 0.004.

Page 3068. In column 1, line 5 from bottom, for "0.65 cal./mole" read "43 cal./mole."—GEORGE SCATCHARD.

1939, VOL. 61

Dorothy Nightingale and Lee Irvin Smith. The Action of Aluminum Chloride on Aromatic Hydrocarbons. I. The 1,3-Dimethyl-4-butylbenzenes.

Page 102. In column 1, the diagram should read:

1,3-dimethyl-4- <i>n</i> -butylbenzene	→	1,3-dimethyl-5- <i>s</i> -butylbenzene
1,3-dimethyl-4- <i>s</i> -butylbenzene	}	→ 1,3-dimethyl-5- <i>t</i> -butylbenzene
1,3-dimethyl-4- <i>t</i> -butylbenzene		

—DOROTHY NIGHTINGALE and LEE IRVIN SMITH.

G. P. Baxter (Chairman), M. Guichard, O. Hönlgschmid and R. Whytlaw-Gray. Ninth Report of the Committee on Atomic Weights of the International Union of Chemistry.

Page 225. In the ninth report of the Committee on Atomic Weights of the International Union of Chemistry through an unfortunate error the uncertainty in the composition of purpureo ruthenium chloride estimated by Gleu and Rehm (*Z. anorg. allgem. Chem.*, 235, 352 (1937)) is given as 0.1–0.2 per cent. instead of 0.01–0.02 per cent., and the uncertainty in the atomic weight of ruthenium as 0.3 unit instead of 0.03 unit. In spite of this optimistic estimate by Gleu and Rehm, the Committee feel that further experimental evidence is necessary before making any change in the atomic weight of ruthenium in the Table.—G. P. BAXTER.

Joseph R. Spies and Thomas H. Harris, Jr. Some Salts of 2-Oxy-6,8-diaminopurine.

Page 352. In column 1, line 1 (after table), for "ninety-three milligrams" read "ninè hundred and thirty milligrams."—JOSEPH R. SPIES.

Harold Hibbert. Studies on Lignin and Related Compounds. XXXVII. The Structure of Lignin and the Nature of Plant Synthesis.

Page 729. In column 2, line 9, for "IV" read "VI."—HAROLD HIBBERT.

Edward S. Amis and Victor K. La Mer. The Entropies and Energies of Activation of Ionic Reactions. The Kinetics of the Alkaline Fading of Brom Phenol Blue in Isoelectric Media.

Page 910. "In Table IV, columns $k_{k=0}$ and Log $k_{k=0}$ for 31.5% EtOH should read 0.971 and -0.013 , respectively; for 42.8% MeOH, the corresponding values are 0.935 and -0.029 ."—VICTOR K. LA MER.

Roger Adams and Madison Hunt. Structure of Gossypol. XIX. Synthesis of 1,2-Dihydroxy-3-isopropyl-6-benzoic Acid.

Page 1132. Column 1, line 2, for "apogossypolic acid (II)" read "apogossypolic acid (V)."—ROGER ADAMS.

Roger Adams and B. R. Baker. Structure of Gossypol. XXI. Synthesis of 1,2-Dimethoxy-3-isopropyl-4-benzoic Acid and of Apogossypolic Acid.

Page 1139. In column 2, line 10 from bottom, read "b. p. 157–158°" for "b. p. 157–168°."

Page 1140. In column 1, line 13 from top, read "added 5 g. of magnesium" for "added 1 g. of magnesium."

Page 1141. In column 1, last line, read " α -bromoacetate" for " α -bromopropionate."—ROGER ADAMS.

F. F. Blicke and H. C. Parke. Alkylaminoalkyl Esters of Aminonaphthoic Acids as Local Anesthetics.

Page 1202. In Table II, the melting point of compound 6 should read "203–204°" instead of "114–115°."—F. F. BLICKE.

Manfred Kiese and A. Baird Hastings. The Dissociation Constant of Hypobromous Acid.

Page 1291. Line 4. "The statement regarding the absence of data on the dissociation constant of hypobromous acid is in error. Shilov and Gladchikova, THIS JOURNAL, 60, 490 (1938), R. M. Chapin, *ibid.*, 56, 2211 (1934), and A. Skrabal and R. Skrabal, *Monatsh.*, 71 273 (1938), have found values for the constant."—MANFRED KIESE and A. BAIRD HASTINGS.

Herbert I. Bernstein and Frank C. Whitmore. The Common Basis of Intramolecular Rearrangements. V. Inversion of Configuration in Semipinacolic Deamination. The Configurational Relationship between (+)-Alanine and (+)-Methylphenylacetic Acid.

Page 1324. Column 1. "The first carbon in the second formula is left electronically deficient by the removal of the amino group with complete octet. This should be indicated by six dots and an asterisk."—F. C. WHITMORE.